**Chemistry Part 1 Review**

1. **Complete the table below.**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Element Name** | **Symbol** | **State at Room Temperature** | **Period #** | **Family Name** | **Metal, Non-metal or Metalloid?** |
| Iron |  |  |  |  |  |
|  | Na |  |  |  |  |
| Beryllium |  |  |  |  |  |
|  | Si |  |  |  |  |

**Multiple Choice – select the best answer below**

1. The first person to organize the elements into the periodic table was:
2. Mendeleev
3. Moseley
4. Bohr
5. Rutherford
6. The valence electrons are:
7. Found in the nucleus
8. The electrons in the orbital closest to the nucleus
9. The electrons in the outermost orbital
10. None of the above
11. The Bohr model of the atom is also known as the:
12. Plum Pudding or Raisin Bun model
13. Planetary Model
14. Nuclear Model
15. Billiard Ball Model
16. Names of elements come from:
17. Scientists
18. Places of discovery
19. Properties
20. All of the above
21. Mendeleev organized the elements by:
    1. Atomic Mass
    2. Atomic Number
    3. Number of neutrons
    4. Alphabetical order
22. The scientist that created the “Nuclear model” of the atom is:
    1. Einstein
    2. Bohr
    3. Rutherford
    4. Thompson
23. The period number of an element tells you:
    1. The number of valence electrons it has
    2. The number of protons it has
    3. The number of orbitals occupied by electrons
    4. The number of electrons it has
24. Who was the scientist that revised Mendeleev’s periodic table so that it was organized by atomic number?
    1. Harry Moseley
    2. Ernest Rutherford
    3. J. J. Thompson
    4. Niels Bohr
25. The nucleus of an atom contains:
    1. Protons and electrons
    2. Protons and neutrons
    3. Neutrons and electrons
    4. Protons, electrons, and neutrons
26. Which of the following can be determined by the atomic number in a neutral atom?
    1. Protons
    2. Neutrons
    3. Electrons
    4. Protons and electrons
27. Which of the following can be determined by the atomic mass?
    1. Protons
    2. Neutrons
    3. Electrons
    4. Protons and electrons
28. The group number of an element tells you:
    1. How many protons it has
    2. How many electrons it has
    3. How many neutrons it has
    4. How many valence electrons it has
29. At the time the first periodic table was created there were roughly:
    1. 38
    2. 60
    3. 92
    4. 118
30. On our most current periodic table, we have how many known elements?
    1. 71
    2. 88
    3. 103
    4. 118
31. **Fill in the information for the following element**:
32. **Draw a neatly labelled Bohr Model for the following element**: